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SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES 67 Thus the solubility is raised to a figure about twice as large as the 6.8×10^{-5} obtained by simply taking the square root of the solubility product [Eq. (3-7)]. At the same time the OH⁻ concentration is lowered slightly, so that the pH is approximately 9.9 instead of 10.0.

SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES

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The solubility product (K_{sp}) is the equilibrium constant of the dissociation reaction of the mineral-containing salt in water. Hence, K_{sp} is a constant at a given temperature. The solubility (S) of a mineral salt is the amount of the salt that is dissolved per unit volume. This quantity may vary, depending on the conditions.

Nutrient Solubility

Mining of coal and metallic minerals in Washington began in the mid to late 1800s. Coal production in Washington hit a high in 2003 with an estimated total production of 6,232,000 short tons, yet production ceased in 2006.

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The book Solutions, Minerals, and Equilibria co-authored in 1965 by Garrels and Charles L. Christ revolutionized aqueous geochemistry. Garrels earned a bachelor's degree in geology from The University of Michigan in 1937. He went on to earn an M.S. degree from Northwestern in 1939, his thesis work was on iron ores of Newfoundland in 1938.

